

# Determination of the Ca/Mg Ratio of Dolomite with the Help of a Pure CO<sub>2</sub> Atmosphere

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## Introduction

Dolomite is a mineral composed of calcium magnesium carbonate, chemically known as CaMg(CO<sub>3</sub>)<sub>2</sub>. It is an essential natural resource with significant technical importance across various industries. One prominent application is in the construction industry, where dolomite is utilized as a building material and an aggregate in concrete. Its hardness and durability make it an ideal component for road base, concrete blocks, and asphalt. Dolomite's resistance to weathering and erosion further enhances its value in infrastructure projects.

Understanding the ratio of calcium to magnesium in dolomite is crucial for optimizing its applications. Varying

ratios of these elements can significantly impact the mineral's physical and chemical properties like solubility, reaction rates and crystalline structure. By knowing and controlling this ratio, manufacturers can tailor dolomite for specific applications, ensuring the desired characteristics and performance.

## Results and Discussion

To investigate the Ca/Mg ratio of dolomite, measurements were carried out with the NETZSCH STA 449 **F3 Jupiter**® using different gas atmospheres. A detailed compilation of the measurement conditions can be found in table 1.

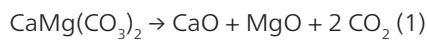
**Table 1** Measurement parameters

Instrument	STA 449 <b>F3 Jupiter</b> ®
Furnace	Silicon carbide furnace
Sensor	TG-DSC Types
Crucible	85 µl Al <sub>2</sub> O <sub>3</sub> with pierced lid
Temperature program	40°C - 1200°C mit at K/min
Atmosphere	70 ml/min of synthetic air or carbon dioxide
Sample mass	approx. 40 mg

## APPLICATIONNOTE Determination of the Ca/Mg Ratio of Dolomite with the Help of a Pure CO<sub>2</sub> Atmosphere

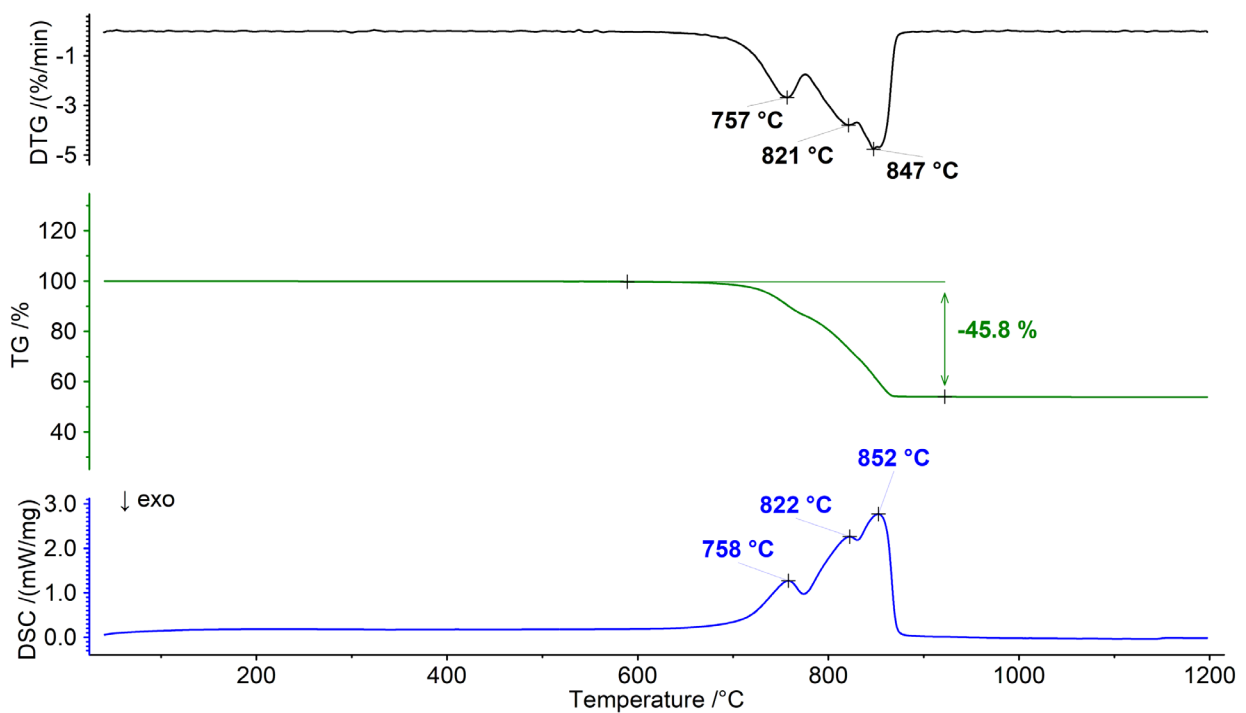
### Thermal Behavior of Dolomite Under Air

When dolomite is heated in the presence of oxygen, it undergoes thermal decomposition. The decomposition reaction can be represented as follows:



During this decomposition process (see figure 1), the solid dolomite breaks down into calcium oxide (CaO),

magnesium oxide (MgO), and carbon dioxide gas (CO<sub>2</sub>). This reaction occurs at temperatures above 700°C and results in a broad mass-loss step which exhibits multiple peaks within the DTG and DSC signal. The overlapping characteristic of the mass loss significantly hinders the ability to assign distinct steps in the TGA curve, thus rendering it impossible to determine any additional information regarding the Ca/Mg ratio of dolomite.



1 Thermal analysis of dolomite under an air atmosphere

### Thermal Behavior of Dolomite Under a Carbon Dioxide Atmosphere

In the presence of a carbon dioxide-containing atmosphere, dolomite exhibits different behavior due to the reversible nature of the decomposition reaction and the varying stability of the carbonates involved. The decomposition reaction (1) can also be represented as two individual equilibrium reactions as follows:

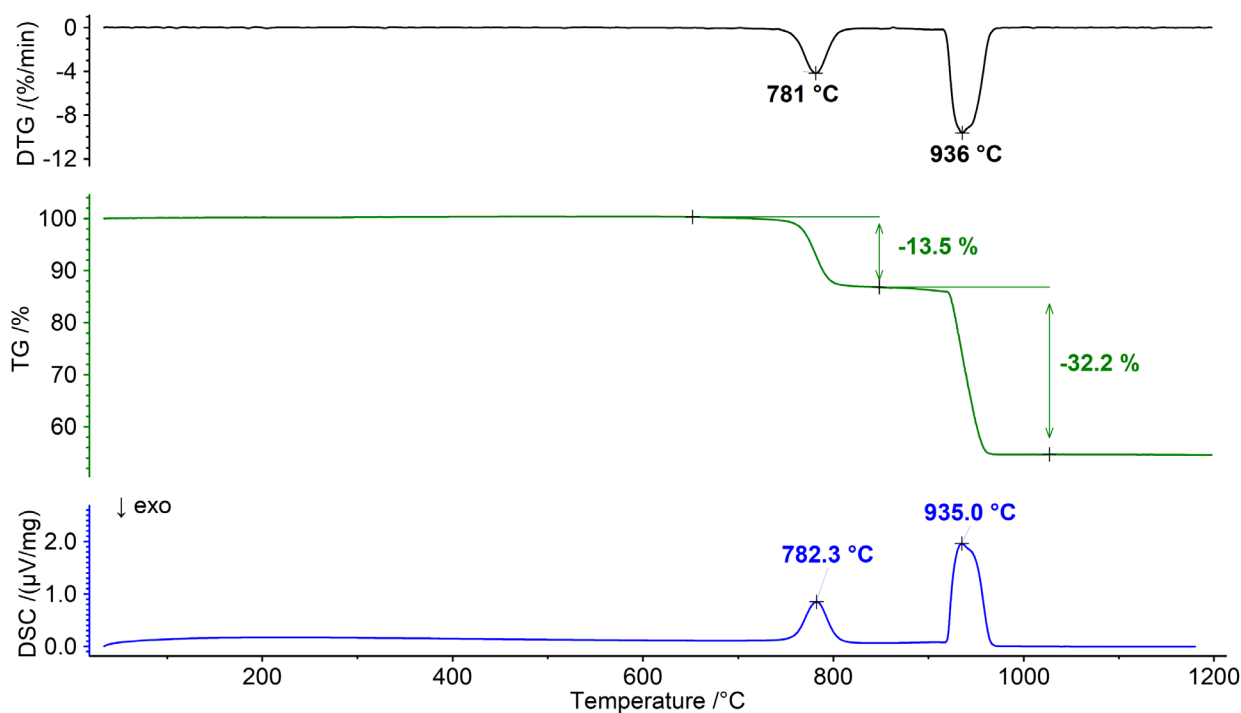


The equilibrium between the reactants and products of the reaction is affected by various factors, including temperature and the concentrations of CO<sub>2</sub>. When carbon dioxide is consistently present, it has an impact on the equilibrium, causing a shift towards the stabilization of carbonates and resulting in a higher decomposition temperature. This effect is more prominent for calcium carbonate than for magnesium carbonate due to its larger

Gibbs Energy of Formation. As a result, the decomposition of calcium carbonate and magnesium carbonate can be clearly distinguished when measured under a pure carbon dioxide atmosphere (see figure 2). This separation offers an opportunity to accurately determine the Ca/Mg ratio of the dolomite, as the initial step can be attributed to the decomposition of MgCO<sub>3</sub> (reaction 2), while the subsequent step corresponds to the decomposition of CaCO<sub>3</sub> (reaction 3).

### Conclusion

Selection of the optimal atmosphere significantly impacts the quality and quantity of information attainable through thermoanalytical measurements. An example of this phenomenon is given in the presented thermal analysis of a dolomite sample. In this case, the transition from a conventional air atmosphere to the slightly more unconventional choice of carbon dioxide resulted in a remarkable impact on the measured thermal properties of the analyzed dolomite. As a consequence, entirely new information like the Ca-Mg ratio became accessible.



**2** Thermal analysis of dolomite under a pure CO<sub>2</sub> atmosphere